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**Knowledge Resource & Relay Centre (KRRC)**


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AIKTC/KRRC/SoP/ACKN/QUES/2021-22/

Date: 12/08/2022School: SoP-PCIBranch: SoPSEM: I

To,  
Exam Controller,  
AIKTC, New Panvel.

Dear Sir/Madam,

Received with thanks the following **Semester/Periodic** question papers from your exam cell:

Sr. No.	Subject Name	Subject Code	Format		No. of Copies
			SC	HC	
1	Human Anatomy and Physiology I	BP101T			
2	Pharmaceutical Analysis I	BP102T		✓	
3	Pharmaceutics I	BP103T		✓	
4	Pharmaceutical Inorganic Chemistry	BP104T			
5	Communication skills	BP105T		✓	
6	Remedial Biology/ Remedial Mathematics	BP106RBT BP106RT		—	

Note: SC – Softcopy, HC - Hardcopy

(Shaheen Ansari)  
Librarian, AIKTC

13 / 7 / 22

End Semester Examination- FH 2022 (Academic Year 2021-2022)

B. Pharm Sem I (Choice based R-2019)

Subject: Pharmaceutical Analysis I

## Q. 1 Attempt all multiple-choice questions (MCQ)

20M

Sr. No.	Questions	Options
1	Number 1500 can have 3 significant figures when it is written as ____.	a $1.50 \times 10^3$
		b $15 \times 100$
		c 1500
		d $150 \times 10$
2	The pH at the equivalence point of a titration of weak base with strong acid is usually ____.	a 5.5
		b 7
		c 8.5
		d 11
3	Weak bases are differentiating solvents for ____.	a Acids
		b Bases
		c Neutral solutions
		d Both acids and bases
4	The concordance between the data and the true value or most probable value is known as ____.	a Stoichiometric end point
		b Standardization
		c Precision
		d Accuracy
5	Starch forms ____ color complex with ____.	a Colorless, sodium thiosulphate
		b Colorless, iodine
		c Blue, iodine
		d Blue, sodium thiosulphate
6	Systematic errors can be removed by ____.	a Buying new instrument
		b Dusting the instrument
		c Breaking the instrument
		d Recalibrating the instrument
7	In direct iodimetry, iodine acts as a mild ____.	a Neutralizing agent
		b Reducing agent
		c Complexing agent
		d Oxidizing agent

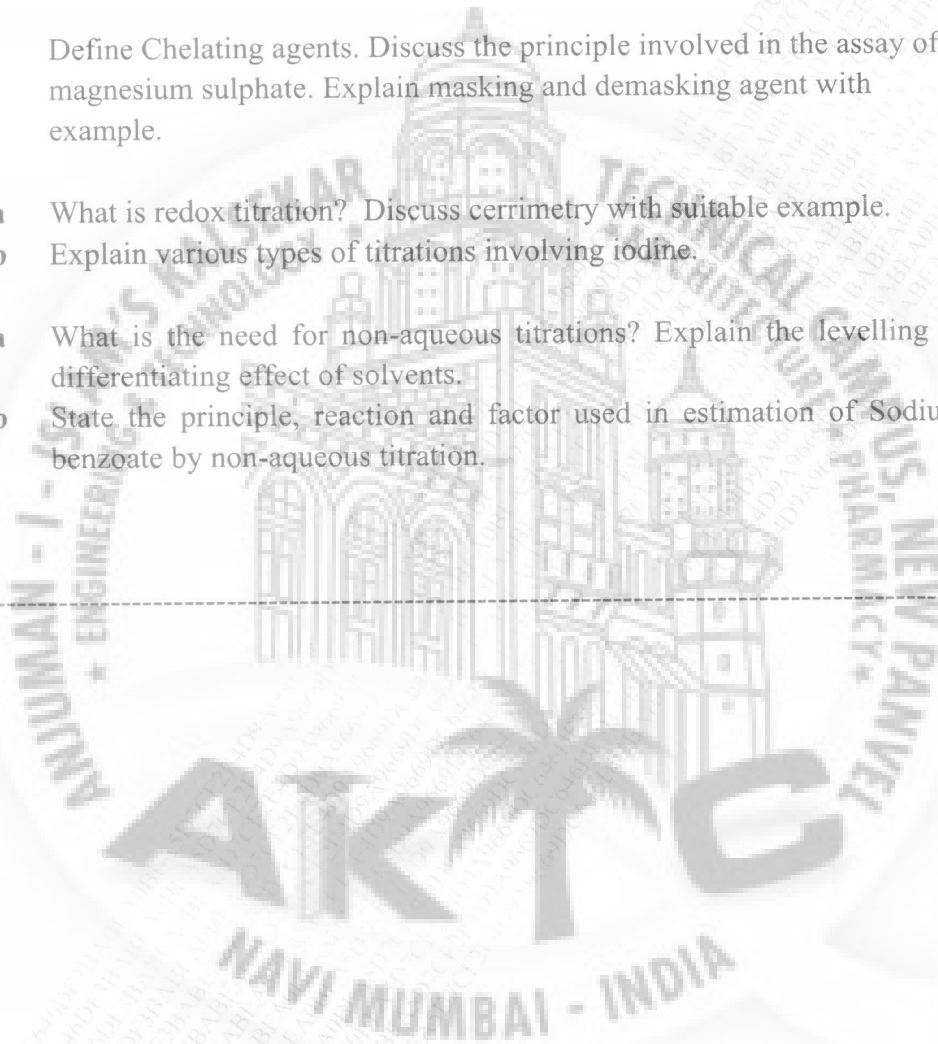
8	_____ is the product of specific conductance and volume of electrolyte.	a	Conductivity
		b	Molar conductance
		c	Conductance enhance
		d	Normal conductance
9	Protogenic solvents are _____ in nature	a	Acidic
		b	Basic
		c	Neutral
		d	Amphoteric
10	Adsorption indicators are used in which method of precipitation titration.	a	Mohr's method
		b	Fajan's method
		c	Gay-Lussac method
		d	Volhard's method
11	The following substances are primary standard EXCEPT _____.	a	Potassium hydrogen phthalate
		b	Anhydrous sodium carbonate
		c	Sodium hydroxide
		d	Arsenic trioxide
12	A substance will precipitate out if the product of its ionic concentrations _____ the Ksp value.	a	Is equal to
		b	Is more than
		c	Is less than
		d	Is double
13	In standardization of perchloric acid, _____ is used as the primary standard.	a	Acetic anhydride
		b	Mercuric acetate
		c	Potassium hydrogen phthalate
		d	Glacial acetic acid
14	In the titration of a strong acid and a weak base, which of the following is used as an indicator?	a	Phenolphthalein
		b	Methyl orange
		c	Potassium dichromate
		d	Alizarin
15	Auxiliary chelating agent or complexing agent that will form complex more strongly with the metal than the titrant under the condition of titration is known as _____.	a	Auxiliary chelating agent
		b	Masking agent
		c	Demasking agent
		d	Complexing agent

16	Quantitative analysis of polarograph is based on _____.	a	Electrode potential
		b	Migration current
		c	limiting current
		d	Half wave potential
17	Indirect titration of iodine is also referred as _____.	a	Iodimetry
		b	Iodometry
		c	Cerrimetry
		d	Permanganometry
18	According to Ostwald theory of indicators, methyl orange in acidic medium is _____ and shows _____ color.	a	Unionized, red
		b	Ionized, red
		c	Unionized, yellow
		d	Ionized, yellow
19	Polarogram of solution containing electro reducible substance is obtained by plotting _____.	a	Current Vs voltage
		b	Current Vs potential
		c	Resistance Vs time
		d	Potential Vs Volume
20	Which of the following analyte can be analyzed by direct complexometric titration?	a	Sodium Chloride
		b	Ammonium chloride
		c	Magnesium sulphate
		d	Aluminum hydroxide

**Q 2. Attempt any one question 12M**

- i. a** Explain the neutralization curve for titration of Strong base and weak acid. **6M**
- b** What is an Electro-analytical cell? Enlist various electrodes used in potentiometry, explain any one electrode in detail **6M**
- ii. a** Discuss in brief the Resonance theory of indicators with suitable example **6M**
- b** Explain the terms half wave potential, diffusion current, limiting current with the help of Polarographic C-V curve. Give the applications of polarography. **6M**

- Q 3. Attempt any four questions** **48M**
- i. Write a brief note on different techniques of analysis. Define and classify error with suitable example. Give the methods for error minimization. **12M**
- ii. a Discuss the theory of fajan's method for precipitation titration. **6M**  
b What are argentometric titrations? Differentiate between Mohr's method and Volhard's method. **6M**
- iii. Define Chelating agents. Discuss the principle involved in the assay of magnesium sulphate. Explain masking and demasking agent with example. **12M**
- iv. a What is redox titration? Discuss cerimetry with suitable example. **6M**  
b Explain various types of titrations involving iodine. **6M**
- v. a What is the need for non-aqueous titrations? Explain the levelling & differentiating effect of solvents. **6M**  
b State the principle, reaction and factor used in estimation of Sodium benzoate by non-aqueous titration. **6M**
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Subject: Pharmaceutics-I

Year and Sem: F.Y. B. Pharm (Sem-I) (2019)

Duration: 3-hour

Total marks: 80

N.B.: 1. All questions are compulsory

2. Figures to right indicate full marks

Q. I	Choose appropriate option for following multiple choice-based questions.	20
1	Use of formulations made up of numerous plants referred as.....	
a	Galenicals	
b	Parenteral	
c	Plant Vehicles	
d	Generics	
2	The first edition of IP was published in ....	
a	1966	
b	1955	
c	1975	
d	1985	
3	The prescription is an order written by a registered medical practitioner to.....	
a	Patient	
b	Pharmacist	
c	Compounder	
d	Nurse	
4	Clarks rule for calculating child's dose is given as	
a	Adult dose x [weight of child (lbs) / 150] = Child's dose	
b	Adult dose x [weight of child (lbs) / 180] = Child's dose	
c	Adult dose x [weight of adult (lbs) / 180] = Child's dose	
d	Adult dose x [weight of child (Kg) / 150] = Child's dose	
5	Powders used for external use are ...	
a	Dusting powder	
b	Bulk powder	
c	Divided powder	
d	Effervescent powders	
6	In liquid dosage form which of the following dosages forms is used for oral administration.	
a	Elixirs	
b	Liniments	
c	Lotion	
d	Enema	
7	_____ are sweet, viscous liquid dosage form of medications meant for cough relief	
a	Linctus	

b	Lozenges	
c	Inhalations	
d	Tablets	
8	Syrup IP contains....%w/w of sugar	
a	66.7	
b	16.6	
c	60.5	
d	56.6	
9	The Proportion of Oil : Water: Gum for preparing primary emulsion of mineral oil by dry gum method is _____	
a	3:02:01	
b	2:02:01	
c	1:01:01	
d	2:01:02	
10	An assembly is used in which a pair of electrodes connected to an electric bulb is dipped into an emulsion and the electric bulb does not glow. Therefore emulsion is .....of type.	
a	o/w	
b	w/o	
c	g/o	
d	o/s	
11	_____ is not a type of semisolid dosage form?	
a	Paste	
b	Creams	
c	Ointments	
d	Suspensions	
12	Suspending agents impart _____	
a	Solubility	
b	Viscosity	
c	Absorption	
d	Wetting	
13	Urethral suppositories also called as _____	
a	Pessaries	
b	Bougies	
c	Enema	
d	Douches	
14	Which form of cocoa butter is considered to be stable for suppository?	
a	Alpha Crystal	
b	Gamma Crystals	

c	Delta Crystals	
d	Beta Crystal	
15	Interaction between two or more substances which may lead to change in color, odor, taste, viscosity and morphology is termed as _____.	
a	Biological incompatibility	
b	Physical incompatibility	
c	Therapeutic incompatibility	
d	Chemical incompatibility	
16	Phase inversion is observed when emulsion prepared by using soaps of monovalent cation is mixed with polyvalent cations. This is example of _____ incompatibility.	
a	Chemical	
b	Physical	
c	Therapeutic	
d	Biological	
17	...cold cream.....is w/o cream	
a	Vanishing cream	
b	Cold cream	
c	Gel	
d	Calamine suspension	
18	Macrogols are used to prepare.....type of ointment base.	
a	Hydrocarbon base	
b	Absorption base	
c	Water miscible base	
d	Water soluble base	
19	.....is an gelling agent	
a	sorbitol	
b	teepol	
c	carbomer	
d	glycerol	
20	The presence of petrolatum-like bases renders them	
a	water miscible	
b	water soluble	
c	Water washable	
d	Occlusive and greasy	
<b>Q. II</b>	<b>ANSWER ANY ONE</b>	<b>(12 Marks)</b>



1. (a)	Write the advantages and disadvantages of powders as the dosage form. Elaborate powders for internal use.	12
2. (a)	Define suspension and discuss formulation of Flocculated and Deflocculated suspension	6
(b)	Differentiate between liniment and lotion. Define enema, elixir and throat paint	6
<b>Q. III</b>	<b>ANSWER ANY FOUR</b>	<b>(48 Marks)</b>
1. (a)	Discuss the brief historical background of the profession of pharmacy in India.	6
(b)	Define dosage form and classify various dosage forms with examples.	6
2. (a)	Enlist different methods of mixing of powders. Explain geometric dilution method.	6
(b)	Explain vehicles and solubilizers use in liquid dosage form	6
3. (a)	Write a short note on suspending agents	6
(b)	Explain instability of emulsions and methods to overcome them	6
4. (a)	Write a note on theobroma oil as a suppository base.	6
(b)	What are pharmaceutical incompatibilities? Give its types and explain physical incompatibility with suitable example	6
5. (a)	Write in brief on excipients in semisolid dosage form	6
(b)	Differentiate between paste and ointment. Define "Gel". Name two natural and two synthetic gelling agents.	6



ANJUMAN-I-ISLAM'S

**KALSEKAR TECHNICAL CAMPUS, NEW PANVEL**Approved by : All India Council for Technical Education, Council of Architecture, Pharmacy Council of India New Delhi,  
Recognised by : Directorate of Technical Education, Govt. of Maharashtra, Affiliated to : University of Mumbai.

- SCHOOL OF ENGINEERING & TECHNOLOGY
- SCHOOL OF PHARMACY
- SCHOOL OF ARCHITECTURE

**SCHOOL OF PHARMACY**

CLASS:- F.Y Pharmacy

SEM: I ATKT

COURSE :- COMMUNICATION SKILLS

DATE: 20/07/2022

DURATION:-1.5Hrs

Max. MARKS:- 35

**Sem 1 ATKT Exam (2021-22)****Questions****Marks**

Questions		Marks
<b>Q 1</b>	<b>Answer the question: ( Attempt any ONE)</b>	<b>10</b>
	a) Define "Communication" and explain the complete cycle of communication along with an example mentioning the elements.  b) What is non-verbal communication. Explain its types with an example.	
<b>Q 2</b>	<b>Answer the question: ( Attempt any FIVE)</b>	<b>25</b>
	a) What is barrier to communication and explain any three types. b) Explain communication as a two-way process. c) What is upward communication? d) Write a note on perspectives in communication? e) How to become an active listener? Explain any three techniques. f) What are the different forms of visual communication? g) What is semantic barrier? Explain with an example.	

\*\*\*ALL THE BEST\*\*\*