



**ANJUMAN-I-ISLAM'S
KALSEKAR TECHNICAL CAMPUS, NEW PANVEL**

Approved by : All India Council for Technical Education, Council of Architecture, Pharmacy Council of India New Delhi,
Recognised by : Directorate of Technical Education, Govt. of Maharashtra, Affiliated to : University of Mumbai.

- SCHOOL OF ENGINEERING & TECHNOL
 SCHOOL OF PHARMACY
 SCHOOL OF ARCHITECTURE

REV:00	QUESTION PAPER SESSIONAL EXAM-02		SEM:- I
CLASS :- First Year B. Pharm			DATE:- 16-03-23
SCHEME:- PCI Syllabus			MARKS:- 30
SUBJECT:- Pharmaceutical Analysis			
DURATION:- 60 mins			
Q.01:		Marks	CO
a) formaldehyde is used as agent a. Masking b. Demasking c. precipitating D. None of the above		1	1
b) In calcium Gluconate assay is used as indicator a. Methyl red c. Erichrome black T b. Alizarin d. Murexide		1	1
c) All of following are reference electrodes except a. Glass electrode c. SHE b. SCE d. Ag/AgCl electrode		1	1
d) Oxidation is..... of electron a. Addition b. Loss c. Gain d. All of the above		1	1
e) NaCl is assayed using 1. AgCl 2. AgNO ₃ 3. AgBr 4. KCl		1	1
f)method is used for Chloride analysis a. Modified Volhard's c. Volhard's b. both d. No		1	1
g) Ceric Ammonium Sulphate is standardized against a. As ₂ O ₃ c. FeSO ₄ b. NaCl d. None of the above		1	1
h) Iodimetry involves titration using a. directly iodine as the titrant c. iodine liberated by reaction b. starch as the titrant d. all f the above		1	1
i) Multidentate ligand EDTA stands for a. Ethylene diamine tetra-acetic acid c. Ethylene diamine tetra-acrylic acid b. Ethane diamine tetra-acetic acid d. Ethylene diacetone tetra-acetic acid		1	1
j) Fajan's method is based onprinciple of precipitation. a. adsorption c. chelation b. Reduction d. All of the above		1	1
Q.02 : Long answers (Any one)		10	2
a) Discuss Mohr's method and estimation of sodium chloride.		10	2
b) Explain Cerimetry.			
Q.03 : Short Answers (Any two)		5	1
a) Explain about hydrogen electrode.		5	2
b) How will you perform analysis of mixture containing 3 metals Cu, Cd and Ca.		5	2
c) Describe calomel electrode.			



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Second Sessional Theory Exam March 2023 DIV A

SUB: Pharmaceutical Inorganic Chemistry

Marks - 30

Time- 3pm to 04 pm (1 hr)

Date – 16/03/2023

	Marks
<p>Q. 01 Multiple Choice Questions (MCQ):</p> <p>1. Povidone is chemically: A. Poly pyrrolidine B. Iodine + Potassium iodide C. Polyvinyl pyrrolidine D. Iodine</p> <p>2. 10% w/v of Iodine and 6% w/v of Potassium Iodide in alcoholic solution is called as a) Strong Iodine Solution b) weak Iodine Solution c) Iodine Tincture d) Aqueous Iodine Solution</p> <p>3. Which one of the following drug is a saline cathartic? A. Magnesium sulphate B. Magnesium trisilicate C. Magnesium carbonate D. Bismuth subcarbonate</p> <p>4. Which of the following radioisotopes has diagnostic application a) Phosphorus-32 b) Chromium-51 c) Cobalt-57 d) All of Above</p> <p>5. Saline cathartic used in barium and lead poisoning is: A. Copper sulphate B. Magnesium sulphate C. Ferrous sulphate D. Sodium sulphate</p> <p>6. Ammonium Chloride is useful in maintaining in acid base equilibrium of the body is also useful as, a) Expectorant. b) Antacid. c) Antioxidant. d) Protective.</p> <p>7. which of the following statement is correct for HCl? a) Help digestion of food b) kills harmful bacteria c) Absorbs Nutrients d) All of the above</p> <p>8. Which of the following is NOT an antacid? a) Calcium Carbonate b) Sodium Bicarbonate c) Kaolin d) Magnesium Oxide</p>	<p>10</p>

<p>9) Which of the following is not water soluble/non systemic antacid?</p> <p>a) Sodium Bicarbonate b) Calcium Carbonate c) Magnesium Carbonate d) Aluminium Hydroxide Gel</p> <p>10) Aqueous Iodine Solution is also known as</p> <p>a) Aluminium Iodide b) Lugol's Iodine c) Weak Iodine d) Iodine Tincture</p>		
<p>Q. 2 Long Questions (Answer ANY 01 out of 02)</p>		10
1.	List the ideal properties of antacids, Discuss the method of preparation, assay, properties and uses of sodium bicarbonate and Aluminium Hydroxide Gel in detail.	10
2.	What are saline cathartics, give classification and explain mechanism of action? Write in detail about Sodium orthophosphate, Bentonite and Kaolin.	10
2.	<p style="text-align: center;">OR</p> <p>Explain in detail different mechanism of antimicrobial agent. Write method of preparation, properties, assay and uses of Potassium Permanganate and Boric Acid.</p>	
<p>Q. 3 Short Questions (Answer ANY 02 out of 03)</p>		10
1.	Explain Iodine and Its preparation and Chlorinated Lime in detail.	5
2.	What are expectorants and Emetics in Detail?	5
3.	<p>Write about</p> <p>a) Precautions to be taken during handling of radiopharmaceutical and storage of Radioactive materials</p> <p>b) Radiopharmaceutical Sodium Iodide I 131</p>	5

*B*est of Luck!!!!!!!



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QUESTION PAPER PERIODIC TEST

CLASS :- First Year B. Pharm (B - Div)

SEM:- I

SCHEME:- PCI Syllabus

SUBJECT:- Pharmaceutical Inorganic Chemistry

DATE:- 17/3/2023

DURATION:- 60 mins

MARKS:- 30

Q.01: MCQs		Marks	CO
a)	The outer layer of teeth is made up of ----- a) Calcium gluconate b) Calcium carbonate c) Calcium bicarbonate d) Calcium chloride	1	
b)	The first edition of IP was published on a) 1956 b) 1965 c) 1955 d) 1985	1	
c)	-----is used as desensitizing agents a) Sodium fluoride b) Calcium carbonate c) Zinc chloride d) Magnesium hydroxide	1	
d)	-----is intracellular electrolyte a) Sodium b) chloride c) bicarbonate d) Potassium	1	
e)	The standard for limit test for iron is made from ----- a) Ferrous sulphate b) Ferric ammonium sulphate c) Ferric sulphate d) Ferrous carbonate	1	
f)	The barium chloride reagent contains. a) BaSO ₄ b) BaCl ₂ & K ₂ SO ₄ c) Sulphate free alcohol d) All	1	
g)	Limit test for lead is also called as ----- test. a) Hydrazine test b) Dithizone test c) Hydrazone test d) Semi carbazone test	1	
h)	90 percent of the cations in the extracellular fluids consist of: a) Hydroxyl ions b) Calcium ions c) Sodium ions d) Bicarbonate ions	1	
i)	Excessive amount of potassium in plasma is termed as ----- a) Hypokalemia b) Hyperpotassaemia c) Hypocalcaemia d) Hypercholeemia	1	
j)	The following is one of the source for Genotoxic impurities a) Catalyst b) By-products c) Intermediate products d) All	1	
Q.02 : Long answers (Any one)			
a)	Define Impurities. Discuss various sources of Impurities.	10	
b)	Define limit test. Explain reaction, principal, procedure for limit test of arsenic with diagram of Gutzeit apparatus or limit test for iron.	10	
Q.03 : Short Answers (Any two)			
a)	Give classification of dental product. What is dental decay? Write a note on anticaries agent.	5	
b)	Explain the importance of electrolytes. Write a short note on extracellular and intracellular electrolytes.	5	
c)	What is monograph? Describe the content of monograph with any one example.	5	

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REV:00	QUESTION PAPER PERIODIC TEST	EXM-04(a)		
CLASS:- First Year B.Pharm		SEM:- I		
SCHEME:- PCI Syllabus				
SUBJECT:-Pharmaceutics-I		DATE:- 17/03/23		
DURATION:- 60 mins		MARKS:- 30M		
Q.01:			Marks	CO
a)	Which of the following are used as aid for handling of diagnostic equipments like Catheters and rectal thermometers? a. Medicated jellies b. Lubricating jellies c. Miscellaneous jellies d. Flavored jellies	1		
b)	_____ are semisolid preparations for external application that differ from similar products in containing a high proportion of finely powdered medicaments. a. Ointment b. Paste c. Gel d. Cream	1		
c)	Which base should be selected when water wash ability is the key requirement? a. A hydrocarbon base b. An absorption base c. An emulsion base d. A water-soluble base	1		
d)	Polyethylene glycol are also known as _____. a. Oleaginous b. Macrogol c. Paraffin d. Lanolin	1		
e)	Which base is likely to be the most occlusive on the skin? a. A hydrocarbon base b. An absorption base c. An emulsion base d. A water-soluble base	1		
f)	Following preparation need to be sipped and swallowed slowly a. Liniment b. Elixir c. Linctus d. Lotions	1		
g)	Following is monophasic liquid dosage form a. Emulsion b. suspension c. syrup d. cream	1		
h)	Cacking is problem in a. Emulsion b. lotions c, deflocculated suspension d. creams	1		
i)	Following are emulsifiers except a. Acacia b. methyl cellulose c. tween d. sucrose	1		
j)	Rheology of biphasic dosage form as a. Thixotropy b. Newtonian c. pseudoplastic d. plastic	1		
Q.02 : Long answers (Any one)				
a)	a. Write a note on water soluble bases used in ointment. b. What is a cream? Differentiate between water in oil and oil in water type of cream.	10		
b)	Write a note on emulsifiers	10		
Q.03: Short Answers (Any two)				
a)	What is a gel? Write a note of different types of gel.	5		
b)	Classify semisolid dosage form. Give the advantages and disadvantages of semisolid dosage form.	5		
c)	Write a note on wetting phenomena in suspension	5		



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REV:00	QUESTION PAPER PERIODIC TEST		EXM-04(a)
CLASS:-First Year B.Pharm			SEM:-I
SCHEME:- PCI Syllabus			
SUBJECT:-Remedial Mathematics			DATE:-18-03-2023
DURATION:- 60 mins			MARKS:-20
Q.1: Solve the Following			
a)	$A = \begin{bmatrix} 1 & 2 & 5 \\ 1 & 1 & 1 \\ 2 & 3 & -1 \end{bmatrix}$ Find A^{-1} by Adjoint method	Marks	CO
b)	$A = \begin{bmatrix} 3 & 2 \\ 7 & 5 \end{bmatrix}$ $B = \begin{bmatrix} 6 & 7 \\ 8 & 9 \end{bmatrix}$ Find $(AB)^{-1}$	05	
c)	If $y = \frac{x+\sin x}{x+\cos x}$ Find $\frac{dy}{dx}$	05	
d)	If $y = \tan x \sin^{-1} x \cos x$ Find $\frac{dy}{dx}$	05	